



## Stoichiometry (Worksheet)

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### Q1.

Given the following reaction:



If it takes 27.4 mL of 0.768 M  $NaOH$  to titrate 16.7 mL of  $H_2SO_4$ , what is the concentration of the  $H_2SO_4$  solution? (hint: balance the equation first)

### Q2.

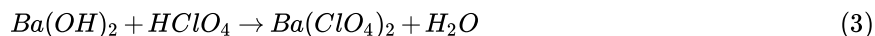
Given the following reaction:



If 24.5 mL of  $HCl$  solution is needed to titrate 33.0 mL of a 0.112 M  $NaOH$ , what is the concentration of the  $HCl$  solution?

### Q3.

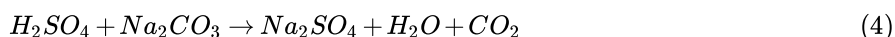
Given the following reaction:



How many mL of 1.2 M  $HClO_4$  is needed to neutralize 5.8 mL of a 0.44 M  $Ba(OH)_2$  solution?

### Q4.

Given the following reaction:



Calculate the molarity of the  $H_2SO_4$  solution if it takes 40.0 mL of  $H_2SO_4$  to neutralize 46.7 mL of a 0.364 M  $Na_2CO_3$  solution.

## Contributors

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